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I. Multiple Choice Questions (Type-I)

1. Which cell will measure standard electrode potential of copper electrode?

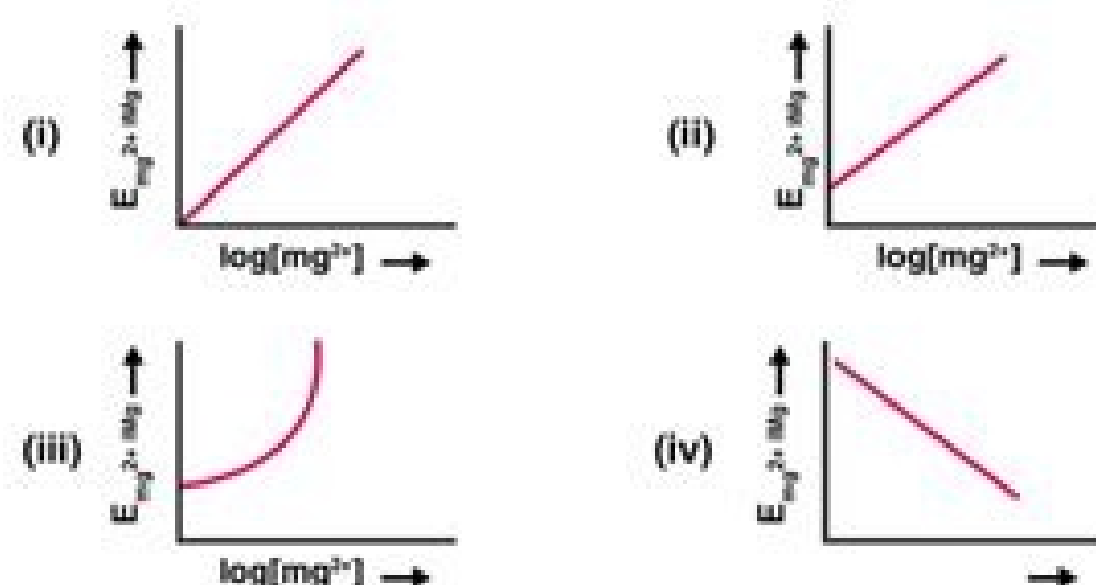
- (i) Pt(s) H₂(g, 0.1 bar) H⁺(aq, 1 M) Cu²⁺(aq, 1 M) Cu
(ii) Pt(s) H₂(g, 1 bar) H⁺(aq, 1 M) Cu²⁺(aq, 2 M) Cu
(iii) Pt(s) H₂(g, 1 bar) H⁺(aq, 1 M) Cu²⁺(aq, 1 M) Cu
(iv) Pt(s) H₂(g, 1 bar) H⁺(aq, 0.1 M) Cu²⁺(aq, 1 M) Cu

Solution:

Option (iii) is the answer.

2. Electrode potential for Mg electrode varies according to the equation

$E_{Mg^{2+}/Mg} = E^{\circ}_{Mg^{2+}/Mg} - 0.0591/2 \log 1/[Mg^{2+}]$. The graph of $E_{Mg^{2+}/Mg}$ Vs $\log [Mg^{2+}]$ is



Solution:

Option (ii) is the answer.

3. Which of the following statement is correct?

- (i) ECell and $\Delta_r G$ of cell reaction both are extensive properties.
(ii) ECell and $\Delta_r G$ of cell reaction both are intensive properties.
(iii) ECell is an intensive property while $\Delta_r G$ of cell reaction is an extensive property.
(iv) ECell is an extensive property while $\Delta_r G$ of cell reaction is an intensive property.

Solution:

Option (iii) is the answer.

4. The difference between the electrode potentials of two electrodes when no current is drawn through the cell is called _____.

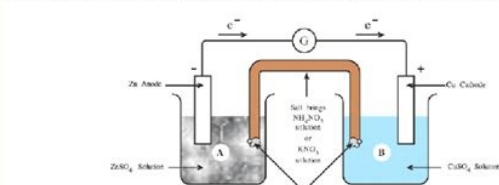
- Cell potential
- Cell emf
- Potential difference
- Cell voltage

Solution:

Option (ii) is the answer.

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The device which transforms the chemical energy into electrical energy is called electrochemical cell. In the electrochemical cell, the energy produced in chemical reaction is transformed into electrical energy.



the salt bridge joins the two solutions and keeps the neutrality of electrical charge

The cell in which the "concentration of the ions associated with the reaction of the cell are 1 mol dm^{-3} and if any gas is associated with the reaction, then its pressure equal to 1 bar and if temperature is 298 K, is called standard cell.

The relative intensity of tendency of receiving the electron is called reduction potential E_{red} or E°_{red} the relative intensity of electrodes for releasing electron is called oxidation potential

 E_{red} or E_{red}° [illegible]

5. Which of the following statement is not correct about an inert electrode in a cell?

- (i) It does not participate in the cell reaction.
(ii) It provides surface either for oxidation or for the reduction reaction.
(iii) It provides a surface for conduction of electrons.
(iv) It provides a surface for a redox reaction.

Solution:

Option (iv) is the answer.

6. An electrochemical cell can behave like an electrolytic cell when _____.

- (i) $E_{\text{cell}} = 0$
(ii) $E_{\text{cell}} < E_{\text{ext}}$
(iii) $E_{\text{ext}} > E_{\text{cell}}$
(iv) $E_{\text{cell}} = E_{\text{ext}}$

Solution:

Option (iii) is the answer.

7. Which of the statements about solutions of electrolytes is not correct?

- (i) The conductivity of the solution depends upon the size of ions.
(ii) Conductivity depends upon the viscosity of the solution.
(iii) Conductivity does not depend upon solvation of ions present in solution.
(iv) The conductivity of the solution increases with temperature.

Solution:

Option (iii) is the answer.

8. Using the data given below to find out the strongest reducing agent.

$E^\circ \text{Cr}_2\text{O}_7^{2-}/\text{Cr}^{3+} = 1.33\text{V}$

$E^\circ \text{MnO}_4^-/\text{Mn}^{2+} = 1.51\text{V}$

$E^\circ \text{Cl}_2/\text{Cl}^- = 1.36\text{V}$

$E^\circ \text{Cr}^{3+}/\text{Cr} = -0.74\text{V}$

(i) Cl^-

(ii) Cr

(iii) Cr^{3+}

(iv) Mn

Solution:

Option (ii) is the answer.

9. Use the data given in Q.8 and find out which of the following is the strongest oxidising agent.

(i) Cl^-

(ii) Mn^{2+}

(iii) MnO_4^-

(iv) Cr^{3+}

Solution:

Option (iii) is the answer.

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No matter how difficult it is, our team is ready to answer all your questions with a guarantee of 100% satisfaction. This chapter contains Class 12 Chemistry multiple choice questions and answers on group 15, 16, 17 and 18 elements, dinitrogen, ammonia, nitrogen oxides, nitric acid, phosphorus and sulphur allotropic forms, sulphuric acid, phosphine, phosphorus halides, phosphorus, sulphur and halogens oxoacids, dioxygen, chlorine, hydrogen chloride and interhalogen compounds. It has an advanced image processing algorithm that can deal with various images of different types. Take a picture of your homework and get answers online With Tutlance homework answers app, you can now take a picture of your homework and get answers online. Once you receive the answers, make sure you rate the tutor as this helps in keeping our marketplace healthy. 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